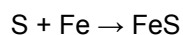
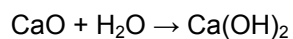
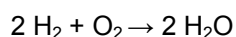
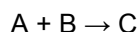


IES Pablo Ruiz Picasso

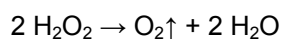
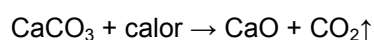
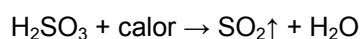
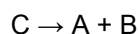


Tipus de reaccions químiques

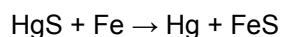
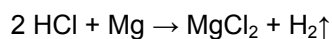
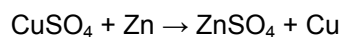
Addició o síntesi. Dos o més substàncies s'uneixen para formar una sola nova substància:



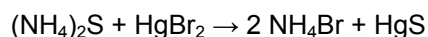
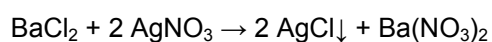
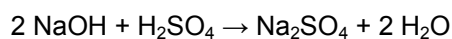
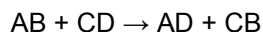
Descomposició o anàlisi. Una sola substància reacciona para formar dos o més noves substàncies:



Simple substitució o simple desplaçament:

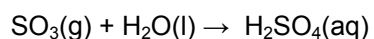
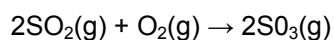
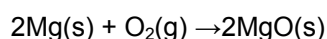
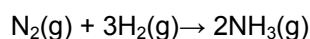
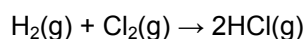


Doble substitució o intercanvi iònic. Dues substàncies reaccionants, intercanvien entre elles llurs ions:

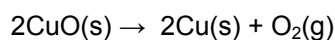
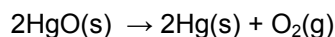


Més exemples de reaccions en química inorgànica

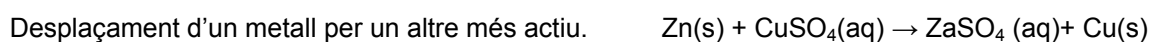
Síntesi:



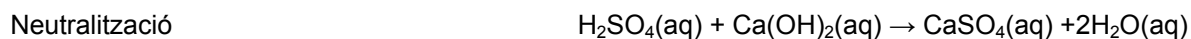
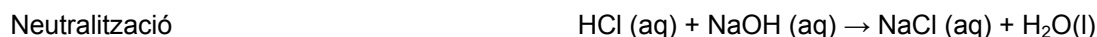
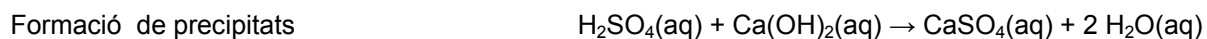
Descomposició:



Substitució o desplaçament:

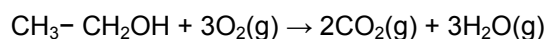
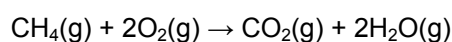


Doble substitució o doble desplaçament:

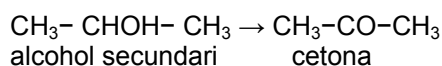
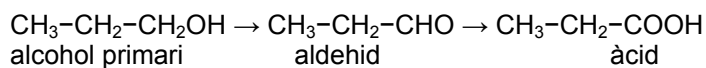


Reaccions orgàniques

Reaccions de combustió:



Reaccions d'oxidació:



Reaccions de substitució:

